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# Water-sensitive multicolor luminescence in lanthanide-organic framework for anti-counterfeiting

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The development of high-level anti-counterfeiting techniques is of great significance in economics and security issues. However, intricate reading methods are required to obtain multi-level information stored in different colors, which greatly limits the application of anti-counterfeiting technology on solving real world problems. Herein, we realize multicolor information anti-counterfeiting under simply external stimulation by utilizing the functional groups and multiple emission centers of lanthanide metal organic framework (Ln-MOFs) to tune luminescence color. Water responsive multicolor luminescence represented by both the tunable color from red to blue within the visible region and high sensitive responsivity has been achieved, owing to the increased nonradiative decay pathways and enhanced  $\text{Eu}^{3+}$ -to-ligand energy back transfer. Remarkably, information hidden in different colors needs to be read with a specific water content, which can be used as an encryption key to ensure the security of the information for high-level anti-counterfeiting.

**Keywords:** metal-organic framework; multicolor luminescence; anti-counterfeiting

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## Introduction

The rampant appearance of counterfeit products has caused tremendous economic losses to customers and copyright owners, and the phenomena from fake medicines to fake military equipment have made it clear that the counterfeit business has become an economic, human health and national security issue<sup>1-3</sup>. In order to solve this issue, anti-counterfeiting techniques have been developed, in which a very effective and convenient method is to put anti-counterfeiting luminescent labels on the products that need to be protected<sup>4-6</sup>. Traditionally, reading out the real information hidden in the fluorescent material requires ultraviolet (UV) excitation. Nev-

ertheless, the unicolor anti-counterfeiting method relies on simple labels produced during use, and its low complexity means that these labels can be easily copied by counterfeiters<sup>7</sup>. Meanwhile, although more complex labels with multicolor and multimodal luminescence have proven to be effective means for high-level anti-counterfeiting, these multicolor or multimodal labels strongly depend on the alternating use of expensive lasers or high-power lasers as the key to read out information<sup>8-12</sup>. This is not an anti-counterfeiting method that can be universally applied to many products on the market. Therefore, a new kind of high-level anti-counterfeiting luminescent materials with multicolor luminescence and

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handy read out model should be explored as an alternative.

As a novel type of multifunctional luminescent materials, metal-organic frameworks (MOFs), constructed by metal ions/clusters and organic ligands through coordinate bonds, have received tremendous attention due to their predictable structures, high porosity, and easily tailorable functions<sup>13–23</sup>. The multi-photon units of MOFs offer it multiple emission centers and tunable energy transmission processes, thus achieving multicolor light emission<sup>24–27</sup>. Especially, the lanthanide metal organic framework (Ln-MOFs), as a subclass of MOFs, can greatly improve the separation of dual-emission peaks from the ligand and metal ions due to the narrow peak emission of lanthanide ions, providing more possibility for the realization of coordinated multicolor changes<sup>28–30</sup>. On the other hand, the tunability of MOFs with organic linkers is of important features due to the unlimited possibility to design multi-functional organic linkers, which can be a point to design the key reading out information<sup>31</sup>. Accordingly, the suitable design and utilization of Ln-MOFs could offer an innovative scheme for the development of multicolor anti-counterfeiting tags under external stimuli.

Considering that chemical anti-counterfeiting methods can be an efficient way for the encryption of information, herein, we present to utilize the multiple centers of MOFs to develop multicolor luminescence under simply external stimuli by precisely designing the synergy emission between metal ions and ligands. The introduction of functional groups as guest interactive sites can interact with external stimuli (the key to analyze information) on the organic ligand, and change the nonradiative decay pathways and energy transfer of MOFs, affecting the final luminescent color. As a proof-of-concept experiment, we design to use the most commonly used water in nature as the key to change the luminescent state of multiple emission centers of a lanthanide MOF, Eu(BDC-NH<sub>2</sub>) (BDC-NH<sub>2</sub> = 2-aminoterephthalate), and thus obtain multicolor luminescence. Benefiting from the suitable stability of Eu(BDC-NH<sub>2</sub>) in water and hydrogen-bond interaction between -NH<sub>2</sub> group and water molecule, as-synthesized Eu(BDC-NH<sub>2</sub>) exhibits gradually enhanced ligand-based luminescence and quenched lanthanide emission with the increasing the water content. And, crucially, the inverse change of emission intensity from ligands and lanthanide ions enables Eu(BDC-NH<sub>2</sub>) to generate colorful luminescence under different water content, and the specific content of water

can be used as the key to add the level of encryption for the anti-counterfeiting ink. Meanwhile, the variable luminescence colors that are visible to the naked eye guarantee the feasibility of Eu(BDC-NH<sub>2</sub>) for *in situ* visualization of the multicolor information.

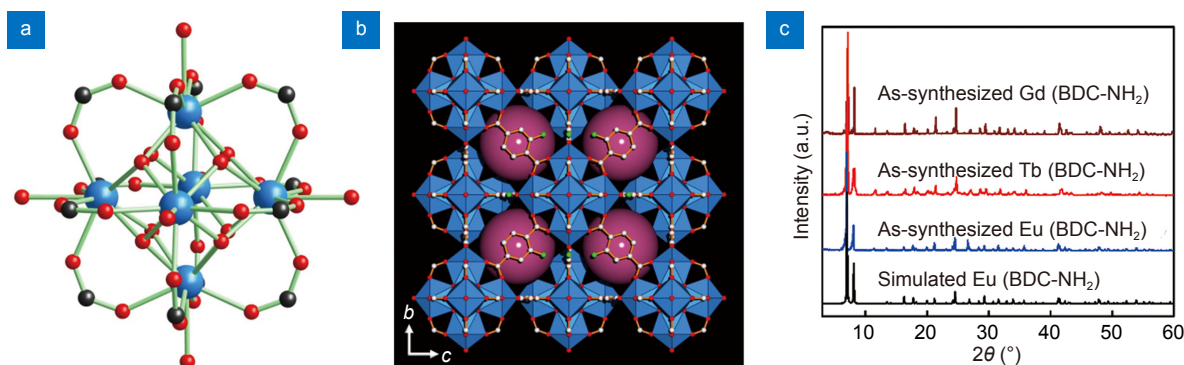
## Experimental section

Synthesis of Eu(BDC-NH<sub>2</sub>), Tb(BDC-NH<sub>2</sub>) and Gd(BDC-NH<sub>2</sub>): Eu(BDC-NH<sub>2</sub>) was synthesized by the solvothermal method. A mixture of Eu(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O (323.4 mg, 0.725 mmol), 2-aminoterephthalic acid (H<sub>2</sub>BDC-NH<sub>2</sub>, 131.3 mg, 0.725 mmol), 2-fluorobenzoic acid (2-FBA, 812.6 mg, 5.8 mmol), DMF (36.5 mL), H<sub>2</sub>O (3 mL), and HNO<sub>3</sub> (2 mL, 3.5 M in DMF) was added to a 100 mL Teflon-lined stainless steel reactor, heated at 110 °C for 60 h, and then slowly cooled to room temperature. The colorless crystals of Eu(BDC-NH<sub>2</sub>) were collected and washed with dimethylformamide (DMF). Elemental analysis: Anal. Calcd. for (Me<sub>2</sub>NH<sub>2</sub>)<sub>2</sub>[Eu<sub>6</sub>(μ<sub>3</sub>-OH)<sub>8</sub>(BDC-NH<sub>2</sub>)<sub>6</sub>(H<sub>2</sub>O)<sub>6</sub>]-5DMF (C<sub>67</sub>H<sub>101</sub>Eu<sub>6</sub>N<sub>13</sub>O<sub>43</sub>, 2688.39): C 29.93, H 3.76, N 6.76; found: C 30.56, H 3.72, N 6.67. Tb(BDC-NH<sub>2</sub>) and Gd(BDC-NH<sub>2</sub>) were synthesized similarly to Eu(BDC-NH<sub>2</sub>) by using Tb(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O and Gd(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O instead of Eu(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O, respectively. X-ray crystal data for Eu(BDC-NH<sub>2</sub>) show: C<sub>67</sub>H<sub>101</sub>Eu<sub>6</sub>N<sub>13</sub>O<sub>43</sub>, *M*<sub>w</sub> = 2688.39, 0.2 × 0.2 × 0.2 mm<sup>3</sup>, Cubic, *Fm* $\bar{3}$ *m*, *a* = *b* = *c* = 21.7579 Å (1 Å = 10<sup>-10</sup> m), *Z* = 4, *T* = 296 K, *F*(000) = 4120, 26525 reflections, 9234 independent reflections, *R*<sub>int</sub> = 0.0444, *R*<sub>1</sub> = 0.0391, *wR*<sub>2</sub> = 0.1082, GOF = 1.006.

[CCDC 1500928 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Center via [www.ccdc.cam.ac.uk/data\\_request/cif](http://www.ccdc.cam.ac.uk/data_request/cif).]

## Results and discussion

The MOF Eu(BDC-NH<sub>2</sub>) was synthesized through the solvothermal reaction between Eu(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O and 2-aminoterephthalic acid (H<sub>2</sub>BDC-NH<sub>2</sub>) in mixed solvents of *N,N*-DMF and water using 2-fluorobenzoic acid (2-FBA) as a cluster directing agent according to the method reported by Eddaoudi et al<sup>32,33</sup>. Eu(BDC-NH<sub>2</sub>) was formulated as (Me<sub>2</sub>NH<sub>2</sub>)<sub>2</sub>[Eu<sub>6</sub>(μ<sub>3</sub>-OH)<sub>8</sub>(BDC-NH<sub>2</sub>)<sub>6</sub>(H<sub>2</sub>O)<sub>6</sub>]-5DMF based on single-crystal X-ray diffraction studies, elemental microanalysis, Fourier transform infrared spectroscopy (FT-IR), and thermogravimetric analysis (TGA) (Figs. S1–S2, Supplementary information). The phase purity of the bulk material was further



**Fig. 1** | (a) Structure of 12-connected hexanuclear Eu cluster  $[\text{Eu}_6(\mu_3\text{-OH})_8(\text{CO}_2)_{12}]$ . Eu, blue; C, black; O, red. (b) Crystal structure of  $\text{Eu}(\text{BDC-NH}_2)$  viewed along the  $a$  axis. Eu, blue polyhedra; C, gray; O, red; N, blue; H atoms,  $\text{Me}_2\text{NH}_2$  cations and free DMF molecules are omitted for clarity. (c) Powder X-ray diffraction patterns of  $\text{Eu}(\text{BDC-NH}_2)$ ,  $\text{Tb}(\text{BDC-NH}_2)$  and  $\text{Gd}(\text{BDC-NH}_2)$ .

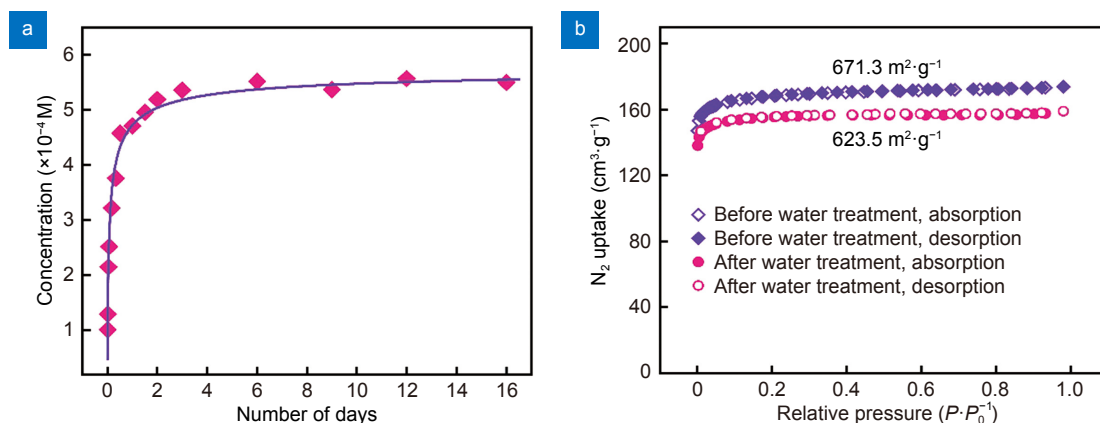
confirmed by similarities between the simulated and measured PXRD patterns (Fig. 1(c)). By replacing the metal salt with  $\text{Tb}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$  or  $\text{Gd}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$  in the same reaction mixture, the isostructural MOFs  $\text{Tb}(\text{BDC-NH}_2)$  and  $\text{Gd}(\text{BDC-NH}_2)$  were obtained as expected, respectively (Fig. 1(c)).

Single-crystal X-ray diffraction analysis indicates that  $\text{Eu}(\text{BDC-NH}_2)$  crystallizes in the cubic  $Fm\bar{3}m$  space group. The crystal structure of  $\text{Eu}(\text{BDC-NH}_2)$  is shown in Fig. 1. In  $\text{Eu}(\text{BDC-NH}_2)$ , each  $\text{Eu}^{3+}$  ion is coordinated by four oxygen atoms from four  $\mu_3\text{-OH}$  groups, four carboxyl oxygen atoms from the four  $\text{BDC-NH}_2^{2-}$  ligands, and one terminal water molecule. The nine-coordinated  $\text{Eu}^{3+}$  ions are further bridged via  $\mu_3\text{-OH}$  and deprotonated carboxylate groups in a bis-monodentate fashion to form hexanuclear europium clusters,  $[\text{Eu}_6(\mu_3\text{-OH})_8(\text{CO}_2)_{12}]$  (Fig. 1(a)). These hexanuclear clusters serve as 12-connected secondary building units (SBUs) and each of them is connected with twelve other clusters through  $\text{BDC-NH}_2^{2-}$  linkers to generate a 3D coordination network with  $fcu$  topology (Fig. 1(b)). The framework structure of  $\text{Eu}(\text{BDC-NH}_2)$  contains octahedral and tetrahedral pores with the diameter of about 7.3 and 4.1 Å, respectively, which are filled with charge balancing cations  $\text{Me}_2\text{NH}_2^+$  and DMF molecules. Nitrogen adsorption measurement performed on the methanol-exchanged  $\text{Eu}(\text{BDC-NH}_2)$  at 77 K shows fully reversible type I isotherms (Fig. S3, Supplementary information), indicative of a porous material with permanent microporosity. The Brunauer-Emmett-Teller (BET) surface area is estimated to be  $671.3 \text{ m}^2 \cdot \text{g}^{-1}$ . Because of the large pore capacity, organic solvent or water molecules can easily access in the pores and the original DMF molecules are exchanged. TGA thermogram of water-exchanged  $\text{Eu}(\text{BDC-NH}_2)$  shows that the maximum water

uptake is about  $0.16 \text{ Kg} \cdot \text{Kg}^{-1}$ .  $\text{Eu}(\text{BDC-NH}_2)$  can retain its crystallinity and unchanged structures in common organic solvents such as THF, ethanol, acetone, acetonitrile and so on, as evidenced by the measured PXRD patterns of soaked  $\text{Eu}(\text{BDC-NH}_2)$  samples in different solvents (Fig. S4, Supplementary information).

The stability of  $\text{Eu}(\text{BDC-NH}_2)$  in water was evaluated by testing the amount of dissolved ligand in supernatant of  $\text{Eu}(\text{BDC-NH}_2)$  aqueous suspension (10 mL,  $10 \text{ mg} \cdot \text{mL}^{-1}$ ) with the elapse of time by UV absorption spectrum. As shown in Fig. 2(a), with the increase of soaking time, the amount of dissolved ligand increased rapidly at the beginning, and the increase rate slowed down after 1 day, and then reached the equilibrium after about 3 days with the equilibrium concentration about  $5.51 \times 10^{-4} \text{ mol} \cdot \text{L}^{-1}$ . According to the molecular formula of  $\text{Eu}(\text{BDC-NH}_2)$ , the dissolved  $\text{Eu}(\text{BDC-NH}_2)$  was calculated to be 2.50% of the original. More interestingly, the undissolved portion of  $\text{Eu}(\text{BDC-NH}_2)$  can maintain its framework and porosity intact. The measured PXRD patterns of  $\text{Eu}(\text{BDC-NH}_2)$  soaked in water for 48 h at room temperature show retained crystallinity and unchanged structures (Fig. S5, Supplementary information). This retention of framework and porosity in water was further authenticated by the preservation of the original apparent BET surface area with a slight decrease, as evidenced by the conserved characteristics of the nitrogen adsorption isotherm (Fig. 2(b)). All of the above results indicate a suitable stability of aqueous  $\text{Eu}(\text{BDC-NH}_2)$  suspension at room temperature.

The excitation spectra of  $\text{Eu}(\text{BDC-NH}_2)$  displays an intense and broad band with a maximum at around 375 nm, which is ascribed to the  $\pi\text{-}\pi^*$  electron transition of  $\text{BDC-NH}_2^{2-}$  ligands (Figs. S6–S9, Supplementary information). Upon excited at 375 nm, the solid sample of



**Fig. 2 |** (a) The concentration of dissolved ligand  $\text{H}_2\text{BDC-NH}_2$  in  $\text{Eu}(\text{BDC-NH}_2)$  suspension at different time. (b)  $\text{N}_2$  sorption isotherms of  $\text{Eu}(\text{BDC-NH}_2)$  before (red) and after (blue) being treated with water at 77 K. Solid symbols: adsorption, open symbols: desorption.

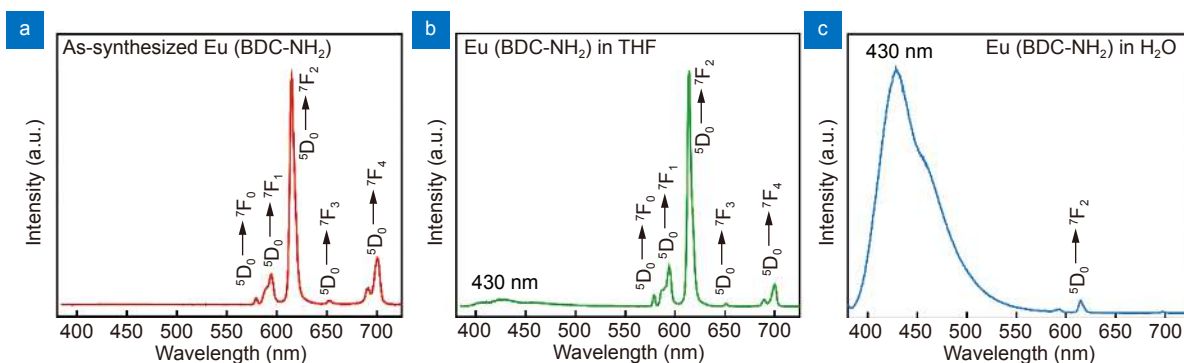
$\text{Eu}(\text{BDC-NH}_2)$  exhibits characteristic emission peaks of  $\text{Eu}^{3+}$  ions at 579, 594, 614, 653, and 700 nm, which can be ascribed to the  $^5\text{D}_0 \rightarrow ^7\text{F}_j$  ( $J = 0, 1, 2, 3$  and 4) transitions, respectively (Fig. 3(a)). It is worth noting that no apparent residual emission from the  $\text{BDC-NH}_2^{2-}$  ligand was observed in solid-state  $\text{Eu}(\text{BDC-NH}_2)$ , indicating that the excitation energy absorbed by the ligands is efficiently transferred to the  $\text{Eu}^{3+}$  ions. Similarly, the suspension of  $\text{Eu}(\text{BDC-NH}_2)$  in THF exhibits the strong characteristic emission peaks of  $\text{Eu}^{3+}$  ions (Fig. 3(b)). However, the characteristic 4f-4f transitions of  $\text{Eu}^{3+}$  ions are almost totally quenched while the ligand-based luminescence centered at 430 nm is drastically enhanced when  $\text{Eu}(\text{BDC-NH}_2)$  is suspended in water (Fig. 3(c)).

Encouraged by the changed luminescence of  $\text{Eu}(\text{BDC-NH}_2)$  under water treatment, we sought to explore its potential for multicolor luminescence through tuning the content of water. As a proof-of-concept experiment, the *in situ* emission spectra were monitored with the piece by piece addition of water in 2 mL of dry organic solvents (THF, ethanol, acetone, and acetonitrile) where  $\text{Eu}(\text{BDC-NH}_2)$  (1 mg) was suspended. As shown in Fig. 4(a, b), the luminescence intensity of  $\text{BDC-NH}_2^{2-}$  lig-

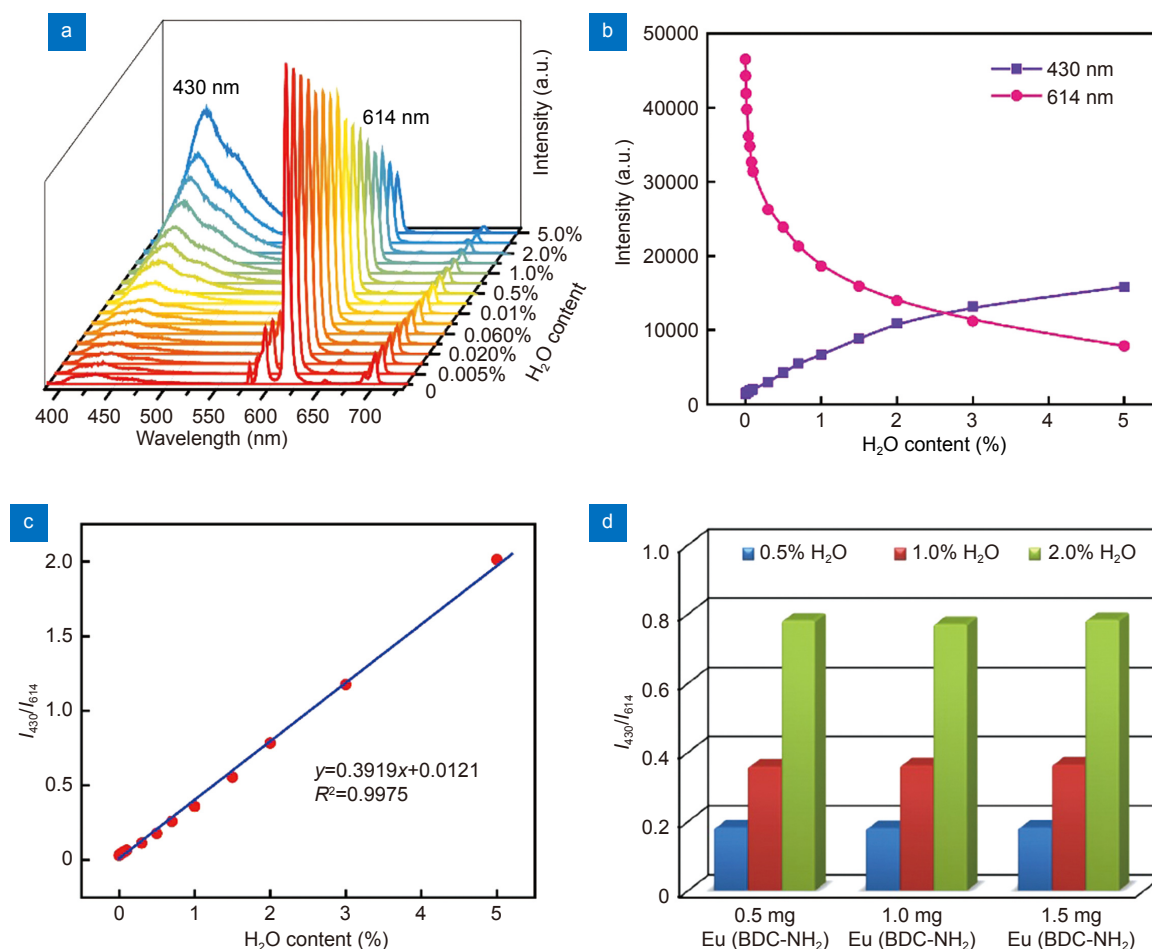
ands at 430 nm substantially increases with the addition of water in THF from 0 to 5 vol%, concomitant with a dramatic decrease of the emissions of  $\text{Eu}^{3+}$  ions when excited at 375 nm. The luminescence intensity ratio between ligand (430 nm) and  $\text{Eu}^{3+}$  ion (614 nm) in  $\text{Eu}(\text{BDC-NH}_2)$  can be correlated well with the water content in THF by a linearly relationship of

$$I_{430}/I_{614} = 0.3919c + 0.0121, \quad (1)$$

with correlation coefficient 0.9975 over a wide range from 0 to 5 vol% (Fig. 4(c)), where  $I_{430}$  and  $I_{614}$  are the luminescence intensity of ligand and  $\text{Eu}^{3+}$ , respectively;  $c$  is the water content in THF (vol%). As shown in Fig. S10 (Supplementary information),  $\text{Eu}(\text{BDC-NH}_2)$  can work even in the ultra-low water content range such as 0–0.1 vol% and 0–1 vol%. Similar trends of emission spectral changes and linear calibration curves were also observed upon adding water to the suspension of  $\text{Eu}(\text{BDC-NH}_2)$  in ethanol, acetone, and acetonitrile (Figs. S11, S12 and S13, Supplementary information). In addition, the fluorescence titrations were also performed using 0.5 and 1.5 mg of  $\text{Eu}(\text{BDC-NH}_2)$  as fluorescent indicator, respectively. The results indicate that the luminescence



**Fig. 3 |** Emission spectra of  $\text{Eu}(\text{BDC-NH}_2)$  in the solid state (a) suspended in THF (b) and in water (c) excited at 375 nm.



**Fig. 4 |** (a) Emission spectra of Eu(BDC-NH<sub>2</sub>) suspended in THF with different water content excited at 375 nm. (b) Intensities of ligand luminescence and <sup>5</sup>D<sub>0</sub> → <sup>7</sup>F<sub>2</sub> transition of Eu(BDC-NH<sub>2</sub>) in the presence of various content of water. (c) The relationship of luminescence intensity ratio ( $I_{430}/I_{614}$ ) and water content. (d) Luminescence intensity ratio ( $I_{430}/I_{614}$ ) measured in THF/H<sub>2</sub>O mixture with given water content (0.5, 1 and 2 vol%) using different amounts of Eu(BDC-NH<sub>2</sub>) (0.5, 1.0 and 1.5 mg).

intensity ratio ( $I_{430}/I_{614}$ ) only depends on the water content while independent of the amount of Eu(BDC-NH<sub>2</sub>) used for multicolor luminescence (Fig. 4(d)), thus demonstrating that Eu(BDC-NH<sub>2</sub>) is an excellent anti-counterfeiting matrix and does not require additional processing steps for reading out multicolor information. To further demonstrate the versatility of this strategy, similar experiments were also performed using the isostructural MOF Tb(BDC-NH<sub>2</sub>) as a sensor. As expected, the luminescence intensity ratio between the ligand and Tb<sup>3+</sup> is also linearly dependent on the water content from 0 to 5 vol% in THF (Fig. S14, Supplementary information).

The multicolor luminescence of designed MOFs Eu(BDC-NH<sub>2</sub>) under external water treatment for multi-level information anti-counterfeiting is feasible, but the stability of Eu(BDC-NH<sub>2</sub>) is also necessary because the anti-counterfeiting tags on the product always face various types of environment. Therefore, it is worth men-

tioning that the photostability of Eu(BDC-NH<sub>2</sub>) is especially outstanding (Fig. S15, Supplementary information) and the fluorescent response toward water is immediate. As is shown in Fig. S16, it takes just 12 s of Eu<sup>3+</sup> emission to accomplish 95% of the overall luminescent intensity variation, and ligand luminescence increases steeply after the addition of water, suggesting the very fast and real-time responsive ability of Eu(BDC-NH<sub>2</sub>). More importantly, the responsive ability of Eu(BDC-NH<sub>2</sub>) can be easily recovered and regenerated by simply removing the dissolved part with organic solvents. As shown by the repeatability experiment (Fig. S17, Supplementary information) and PXRD (Fig. S18, Supplementary information), the frameworks of Eu(BDC-NH<sub>2</sub>) are also intact and the luminescent intensity ratio is fully reversible without significant hysteresis after four consecutive of detection cycles.

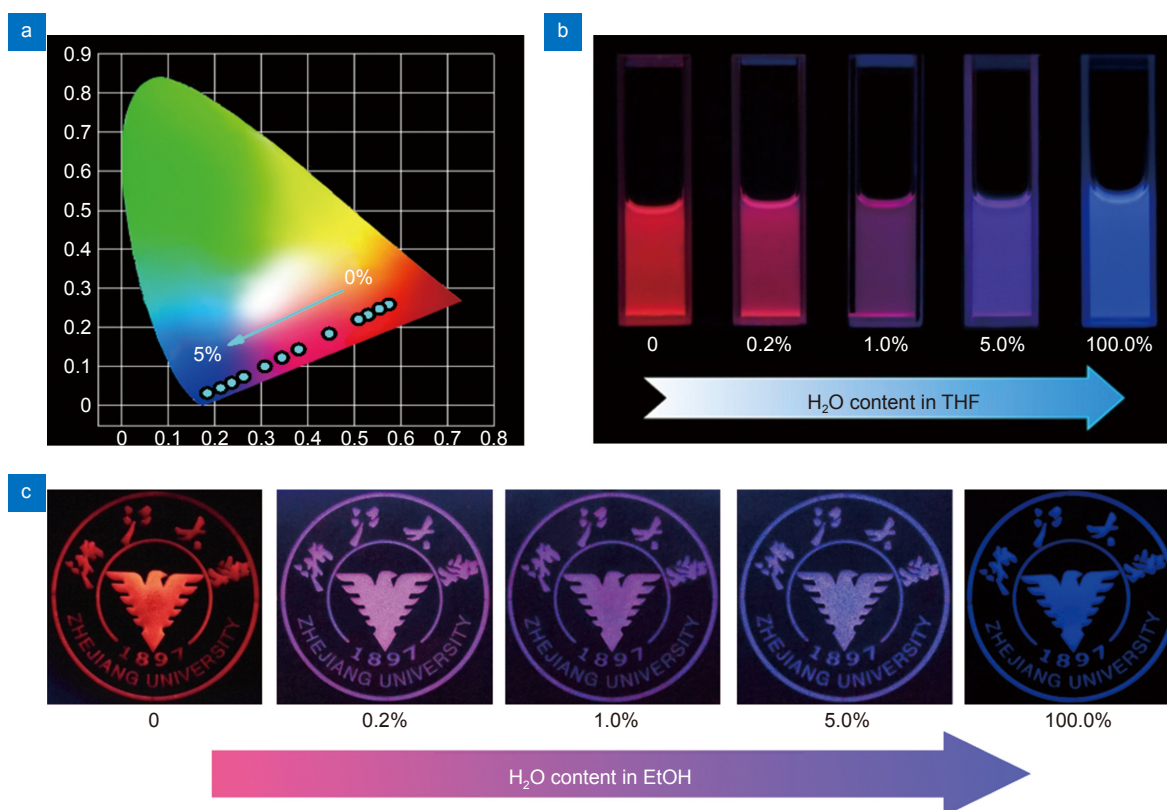
Due to the different luminescent colors and opposite

intensity changes of ligand and  $\text{Eu}^{3+}$  ion, the  $\text{Eu}(\text{BDC-NH}_2)$  sensor exhibits a prominent luminescence color change, which is almost linearly tuned from red to blue with the addition of water in THF. Figure 5(a) shows the emission colors of  $\text{Eu}(\text{BDC-NH}_2)$  in Commission Internationale de L'Eclairage (CIE) chromaticity diagram transformed from the corresponding water-dependent emission spectra, which changes from (0.562, 0.295) to (0.184, 0.080) when the water content increases from 0 vol% to 5 vol%. It can be found that the color evolution strongly depends on the water content and even a slight increase of the water content can cause obvious color change in THF and EtOH (Fig. 5(b, c)). Especially, different water content is used as a key to read anti-counterfeiting information under specific color, which greatly improves the security of anti-counterfeiting materials. Meanwhile, such a significant color switching can be easily observed by the naked eye and captured by a single CCD camera, thus enabling us to directly visualize the water content in various organic solvents. Above results demonstrate that  $\text{Eu}(\text{BDC-NH}_2)$  is an excellent candidate for anti-counterfeiting materials with real-time, robust and reusable multicolor luminescence by simple water treatment, thus providing a significant advantage

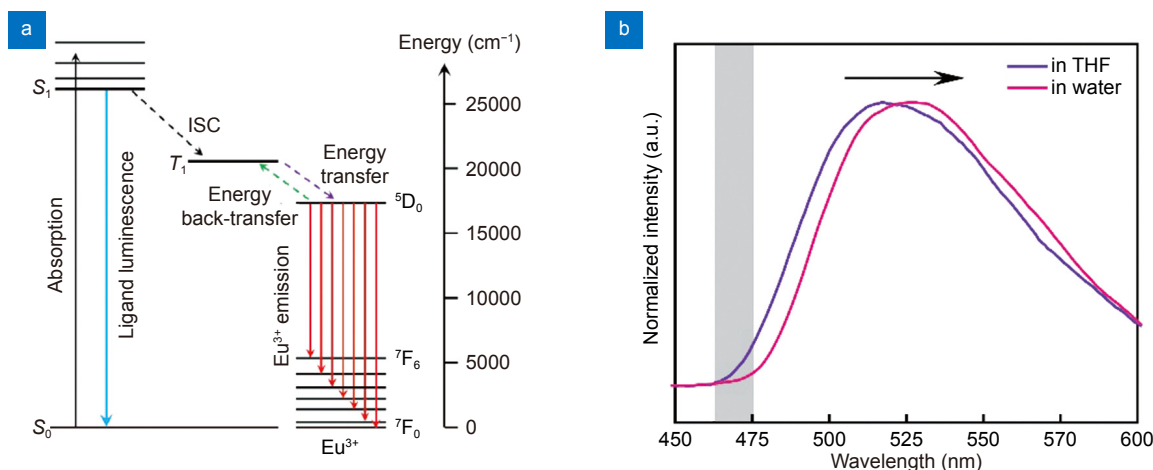
over the traditional anti-counterfeiting tags.

Such a highly sensitive luminescence response of  $\text{Ln}(\text{BDC-NH}_2)$  toward water is mainly due to the following three reasons: 1) hydrogen-bond interaction between  $-\text{NH}_2$  group and water molecules; 2) quenching of  $\text{Ln}^{3+}$  emission by the effect of O-H oscillators, which is universal and acknowledged and 3) the changed energy transfer process of  $\text{Ln}(\text{BDC-NH}_2)$ .

Among these reasons, the quenching of  $\text{Ln}^{3+}$  emission caused by O-H oscillators is the dominant one. This is because  $\text{BDC-NH}_2$  possesses hydrophilic active sites  $-\text{NH}_2$ , thus water molecules enter cages easily, and meanwhile, the introduction of water results in the increase of nonradiative decay pathways (O-H oscillators) that prevents radiative relaxation (emission), accounting for the quenching of  $\text{Ln}^{3+}$  emission<sup>34,35</sup>. On the other hand, we speculate that there may be another mechanism. The energy levels of the ligands and  $\text{Eu}^{3+}$  ions of  $\text{Eu}(\text{BDC-NH}_2)$  with water present and absent were further investigated. According to antenna effect theory, the energy level of triplet state ( $T_1$ ) of ligand has to be at about 2500–4000  $\text{cm}^{-1}$  above the emissive  $4f$  level of  $\text{Eu}^{3+}$  ( $^5D_0$ , 17300  $\text{cm}^{-1}$ ) in order to efficiently sensitize  $\text{Eu}^{3+}$  emission, otherwise the energy back-transfer becomes



**Fig. 5 |** (a) CIE chromaticity coordinates of the luminescence color of  $\text{Eu}(\text{BDC-NH}_2)$  in THF with different water content ( $c = 0, 0.01, 0.04, 0.1, 0.3, 0.5, 0.7, 1, 1.5, 2, 3,$  and  $5$  vol%, respectively). Photographs of  $\text{Eu}(\text{BDC-NH}_2)$  in (b) THF and (c) EtOH with different water content excited at 365 nm.



**Fig. 6 |** (a) Schematic representation of energy transfer process in Eu(BDC-NH<sub>2</sub>). (b) Phosphorescence spectra of Gd(BDC-NH<sub>2</sub>) at 77 K excited at 375 nm in frozen THF and water.

possible and the remaining ligand fluorescence in combination with the Eu<sup>3+</sup>-centered emission can be simultaneously observed (Fig. 6(a))<sup>36,37</sup>. The lowest energy level of Gd<sup>3+</sup> (32200 cm<sup>-1</sup>) is well above the S<sub>1</sub> and T<sub>1</sub> levels of BDC-NH<sub>2</sub>, and therefore, the triplet energy level of the corresponding ligands can be obtained from the phosphorescence spectra of the Gd(BDC-NH<sub>2</sub>). Based on the phosphorescence spectra of Gd(BDC-NH<sub>2</sub>) at 77 K in frozen solvents, the T<sub>1</sub> level of BDC-NH<sub>2</sub><sup>2-</sup> ligand in THF was calculated to be 21277 cm<sup>-1</sup> while decreased to 20833 cm<sup>-1</sup> in water (Fig. 6(b)), which may be attributed to the hydrogen-bond interaction between -NH<sub>2</sub> group of BDC-NH<sub>2</sub><sup>2-</sup> ligands and water molecules. Consequently, upon increasing the water content in THF, the hydrogen-bond interaction between -NH<sub>2</sub> group and water is promoted and the BDC-NH<sub>2</sub><sup>2-</sup> ligand with low T<sub>1</sub> level is dominant, thus enhancing the energy back-transfer probability from Eu<sup>3+</sup> ion to ligand. As a result, the opposite intensity variation of Eu<sup>3+</sup> emission and ligand emission make the intensity ratio more sensitive to water, thus enhancing the detection efficiency. The fluorescence lifetimes of Eu(BDC-NH<sub>2</sub>) in THF and EtOH with different water content were monitored at 430 and 614 nm. The lifetime of <sup>5</sup>D<sub>0</sub> (Eu<sup>3+</sup>) is reduced by the addition of water, whereas the lifetime of ligand-based emission at 430 nm increases (Figs. S19 and S20, Supplementary information). This is well consistent with the proposed energy back-transfer mechanism.

## Conclusion

In summary, we have successfully illustrated a proof of concept for real-time and multicolor luminescence under simple water treatment in organic solvents using a lanthanide MOF Eu(BDC-NH<sub>2</sub>). Due to the unique re-

sponsive mechanism based on increased nonradiative decay pathways and Eu<sup>3+</sup>-to-ligand energy back transfer process, the ligand-based luminescence in Eu(BDC-NH<sub>2</sub>) significantly enhanced while the Eu<sup>3+</sup> emission decreased with the addition of water content. Accordingly, the synergistic luminescence of Eu(BDC-NH<sub>2</sub>) from ligand and Eu<sup>3+</sup> emissive centers exhibits a broad span range in CIE chromaticity coordinates, resulting in multicolor luminescence in visible range. More importantly, the different emission colors obtained by water treatment correspond to specific water content, which can be used as an encryption key for high-level anti-counterfeiting tags. Meanwhile, the variable luminescence color makes Eu(BDC-NH<sub>2</sub>) have excellent potential for *in situ* visualizing the multicolor information in organic solvents straightforwardly. The strategy we report here would open a new avenue to develop multicolor luminescent MOFs under simple external stimuli for high-level anti-counterfeiting and stimulate more applied research.

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## Competing interests

The authors declare no competing financial interests.

## Supplementary information

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